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n = number of molecules, R . T = Boltzmann's constant, T = temperature [in K] • 2. nd. form : $p = \rho RT$ = Engineering Thermodynamics Equation Sheet Heat & thermodynamics Cheat sheet. Saved by Grace Rapai. 53. Electrical Engineering Quotes Engineering Humor Engineering Science Chemical Engineering Engineering Projects Engineering Notes Physics Humor Engineering Technology Electronic Engineering. Heat & thermodynamics Cheat sheet | Physics classroom ... Thermodynamics key facts (3/9) • Heat energy needed to change phase • The latent heat L determines the energy needed to change the phase of a mass m . • Units of L will be J/kg. -1. - can be fusion or vaporization • This energy is either absorbed (solid → liquid → gas) or released (gas → liquid → solid) • A phase change takes place at constant temperature. Revision : Thermodynamics level 1. pugsly56. 2 points · 10 months ago. Lucky. My professor didn't let us use anything in our test other than a pencil. Calculator and the thermodynamic property table book. No cheat sheet, reference sheet. Nothing. Thermodynamics 2 Reference

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Thermodynamics $\dot{Q} = k A \frac{\Delta T}{L}$
 \dot{Q} = rate of heat transfer
 Q = thermal energy
 A = Area of thermal conductivity
 U = coefficient of heat conductivity (U-factor)
 ΔT = change in temperature
 R = resistance to heat flow (R-value)
 $\dot{P} = \sigma A T^4$
 \dot{P} = net power radiated = $5.6696 \times 10^{-8} T^4$
 T_1, T_2 = temperature at time 1, time 2
 v = flow velocity
 Engineering Formula Sheet IV. The First Law of Thermodynamics as a Rate Equation A. Most general form $dE_{c.v.}/dt = \dot{Q}_{c.v.} - \dot{W}_{c.v.} + \dot{m}_{in} e_{in} - \dot{m}_{out} e_{out}$
 $dE_{c.v.}/dt$ = rate of change of energy in c.v.
 $\dot{Q}_{c.v.}$ = rate of heat added to c.v.
 $\dot{W}_{c.v.}$ = rate of work done + rate of energy flow in to c.v. - rate of energy flow out of c.v.
 B. For a steady flow process $dE_{c.v.}/dt = 0$ and $\dot{m}_{in} = \dot{m}_{out} = \dot{m}$
 $\dot{Q}_{c.v.} - \dot{W}_{c.v.}$ = $\dot{m}(e_{out} - e_{in})$
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Thermodynamics key facts (3/9) • Heat energy needed to change phase • The

latent heat h_{fg} determines the energy needed to change the phase of a mass m .

- Units of h_{fg} will be kJ/kg . -1 . - can be fusion or vaporization
- This energy is either absorbed (solid \rightarrow liquid \rightarrow gas) or released (gas \rightarrow liquid \rightarrow solid)
- A phase change takes place at constant temperature.

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Thermodynamics: the study of energy, energy transformations and its relation to matter. The analysis of thermal systems is achieved through the application of the governing conservation equations, namely Conservation of Mass, Conservation of Energy (1st law of thermodynamics), the 2nd law of thermodynamics and the property relations.

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IV. The First Law of Thermodynamics as a Rate Equation A. Most general form $dE_{c.v.}/dt = \dot{Q}_{c.v.} - \dot{W}_{c.v.} + \dot{m}_{in}e_{in} - \dot{m}_{out}e_{out}$ rate of change of energy in c.v. = rate of heat added to c.v. rate of work done + rate of energy flow in to c.v. rate of energy flow out of c.v. B. For a steady flow process $d/dt = 0$ and $\dot{m}_{in} = \dot{m}_{out} = \dot{m}$ $\dot{Q}_{c.v.} - \dot{W}_{c.v.}$